

antwoorden opgaven

- 1 a gas
b vrij te kiezen
- 2 a vloeistof
b nee, de verhouding ligt vast
- 3 a 2 N 4 H 3 O
b 3 Ca 2 P 8 O
c 6 C 12 H 6 O
d 3 K 1 Fe 6 C 6 N
e 1 Cu 4 N 12 H 1 S 4 O
f 2 K 2 Cr 7 O
- 4 a 10 H 5 O
b 10 C 20 O
c 300 C 600 H 100 O
d 4 C 28 H 14 O
e 120 C 240 O 20 H 100 Na
f 5 Fe 10 N 100 H 10 S 70 O
- 5 a $\text{C}_3\text{H}_8\text{O}_2 + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
b $\text{P}_4 + 5 \text{O}_2 + 6 \text{H}_2\text{O} \rightarrow 4 \text{H}_3\text{PO}_4$
c $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
d $2 \text{Fe} + 3 \text{Cl}_2 \rightarrow 2 \text{FeCl}_3$
e $4 \text{FeS} + 7 \text{O}_2 \rightarrow 2 \text{Fe}_2\text{O}_3 + 4 \text{SO}_2$
f $\text{Br}_2 + 2 \text{FeBr}_2 \rightarrow 2 \text{FeBr}_3$
g $2 \text{SnCl}_2 + \text{O}_2 + 4 \text{HCl} \rightarrow 2 \text{SnCl}_4 + 2 \text{H}_2\text{O}$
h $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 6 \text{H}_2\text{O} + 4 \text{CO}_2$
i $3 \text{SO}_3 + \text{Fe}_2\text{O}_3 \rightarrow \text{Fe}_2(\text{SO}_4)_3$
j $\text{SiO}_2 + 4 \text{HF} \rightarrow \text{SiF}_4 + 2 \text{H}_2\text{O}$
k $6 \text{HCl} + \text{Al}_2\text{O}_3 \rightarrow 2 \text{AlCl}_3 + 3 \text{H}_2\text{O}$
l $4 \text{HNO}_3 \rightarrow 2 \text{H}_2\text{O} + 4 \text{NO}_2 + \text{O}_2$
m $(\text{NH}_4)_2\text{Cr}_2\text{O}_7 \rightarrow \text{Cr}_2\text{O}_3 + 4 \text{H}_2\text{O} + \text{N}_2$
- 6 2 S₈: 2 moleculen die ieder bestaan uit 8 atomen zwavel
8 S₂: 8 moleculen die ieder bestaan uit 2 atomen zwavel
16 S: 16 atomen zwavel
- 7 a waterstof, zuurstof, stikstof, chloor, broom, fluor
b kwik
- 8 AlCl₃(s): verhoudingsformule (vaste stof)
AlCl₃(g): molecuulformule (moleculen in gasvorm)